

CHAPTER NO 1

SHORT QUESTIONS

Q No.1 Define organic chemistry.

Q No.2 Differentiate between atomic number and atomic mass.

Q No.3 What is empirical formula?

Q No.4 Define ion, cation, anion and free radical.

Q No.5 What is Avogadro number.

Q No.6 What do you know about corpuscular nature of matter?

LONG QUESTIONS

Q No.1 Note on Chemical formula.

Q No.2 See all numerical of chapter 1

CHAPTER NO 2

SHORT QUESTIONS

Q No.1 Write defects of Rutherford atomic model.

Q No.2 Define isotopes.

Q No.3 Write daily uses of isotopes.

Q No.4 Give electronic configuration of oxygen.

Q No.5 Give isotopes of uranium.

Q No.6 Write isotopes of Carbon.

Q No.7 Write conclusion of Rutherford atomic model.

LONG QUESTIONS

Q No.1 Note of Rutherford atomic model.

Q No.2 Note on Bohr atomic model.

Q No.3 Note on isotopes of carbon.

Q No.4 See Exercise of chapter No 2

CHAPTER NO 3

SHORT QUESTIONS

- Q No.1 Define group and period of periodic table.
- Q No.2 Differentiate between alkali metals and alkaline earth metal.
- Q No.3 Differentiate between noble gases and zero group.
- Q No.4 How many periods are present in periodic table.
- Q No.5 How many groups are present in periodic table.
- Q No.6 What is shielding effect.
- Q No.7 Define atomic size.
- Q No.8 Define electronegativity.

LONG QUESTIONS

- Q No.1 Note on electron affinity.
- Q No.2 Note on ionization energy.
- Q No.3 Note on atomic size.
- Q No.4 Brief about division of periodic table.

CHAPTER NO 4

SHORT QUESTIONS

Q No.1 What is octet rule?

Q No.2 What is duplet rule?

Q No.3 What is chemical force or chemical bond?

Q No.4 Define ionic bond.

Q No.5 What is covalent bond.

Q No.6 Define hydrogen bonding.

Q No.7 Give properties of ionic bond.

Q No.8 What is triple covalent bond. Give its example.

LONG QUESTIONS

Q No.1 Note on ionic bond.

Q No.2 Note on covalent bond.

Q No.3 Note on hydrogen bonding.

Q No.4 Write applications of hydrogen bonding.

CHAPTER NO 5

SHORT QUESTIONS

Q No.1 Define diffusion

Q No.2 Define Boyls law

Q No.3 Why evaporation cause cooling?

Q No.4 Define Charls law

Q No.5 Why boiling point is decreases on hight?

Q No.6 What is sublimation?

Q No.7 Differentiate between crystalline solid and amorphous solid.

Q No.8 What is allotropy give its example.

LOND QUESTIONS

Q No.1 Note on Bolys law

Q No.2 Note on Charls law

Q No.3 Note on allotropy.

CHAPTER NO 6

SHORT QUESTIONS

Q No.1 Gasoline does not dissolve in water why?

Q No.2 Define supersaturated solution.

Q No.3 Give the components of steel alloy.

Q No.4 What is colloids? Give its properties.

Q No.5 What is suspension? Give its properties.

Q No.6 What is molarity?

Q No.7 Define solution..

Q No.8 How we can prepare urea.

LOND QUESTIONS

Q No.1 Define saturated, unsaturated, and supersaturated solution.

Q No.2 Note on solution of solid.

Q No.3 Note on solution of liquid.

Q No.4 Note on solute solvent interaction.

Q No.5 How temperature effect on solubility?

CHAPTER NO 7

SHORT QUESTIONS

Q No.1 Find oxidation number of Cr in $K_2Cr_2O_7$.

Q No.2 Define oxidation.

Q No.3 Differentiate between oxidizing agent and reducing agent?

Q No.4 What is galvanic cell?

Q No.5 Define dry cell.

Q No.6 What is tinning?

Q No.7 Define reduction.

Q No.8 What is electroplating?

Q No.9 What is cathodic protection?

Q No.10 What is galvanizing?

Q No.11 What is corrosion?

Q No.12 Brief note on rusting of iron.

LOND QUESTIONS

Q No.1 Note on electrochemical cell.

Q No.2 How we prepare sodium metal from fused sodium chloride.

Q No.3 Prepare of sodium hydroxide from brine.

Q No.4 Note on electroplating on steel.

Q No.5 How we prevent from corrosion?

Q No.6 Note on galvanic cell.

Q No.7 How we can give oxidation number or state to an element?